

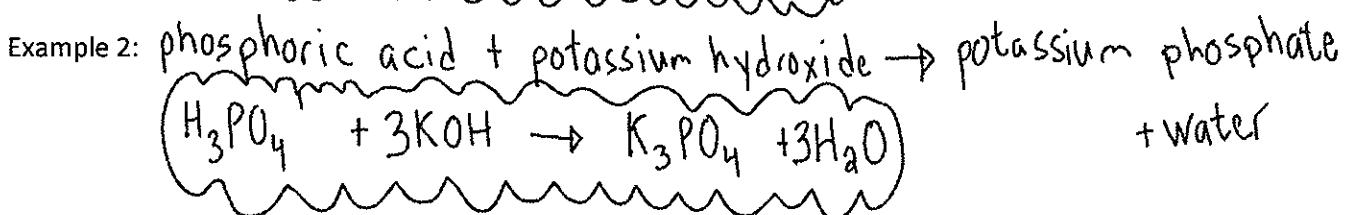
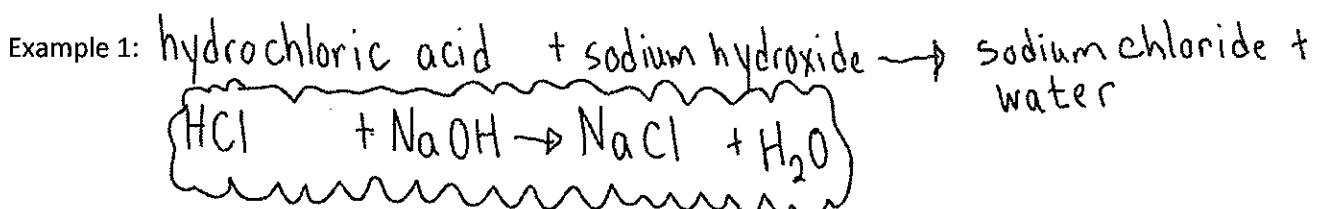
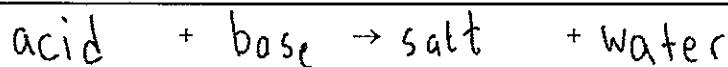
Acid-Base Reactions

Science 10

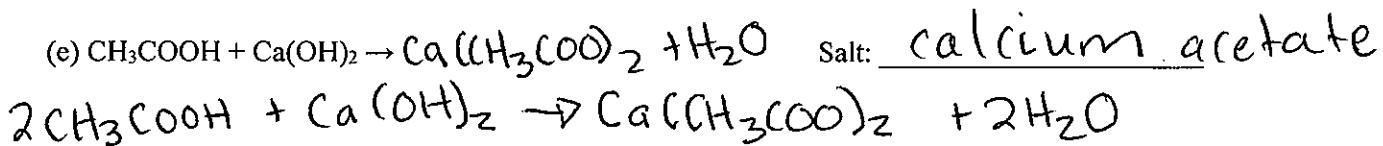
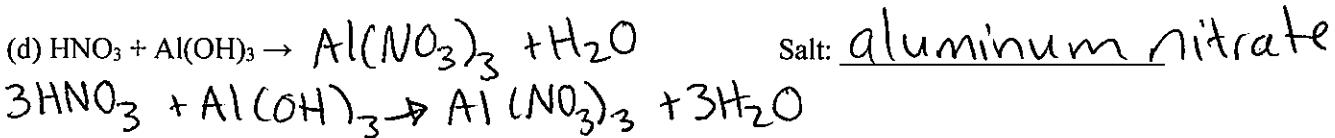
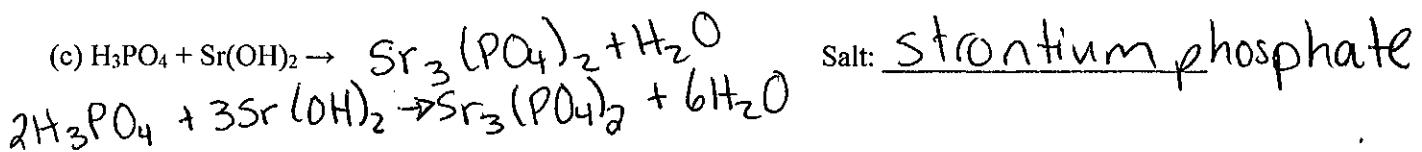
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Neutralization (acid-base): when an acid reacts with a base to form a salt and water



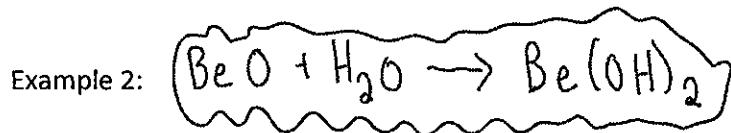
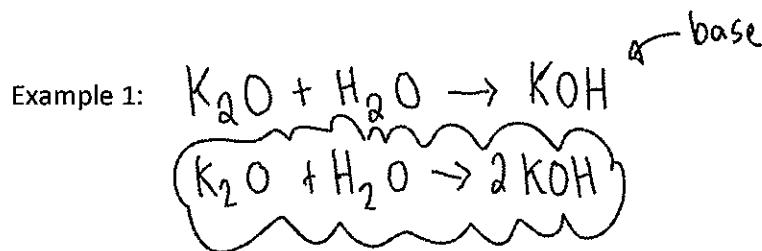
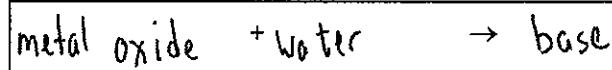
Complete and balance the following equations. Write the name of the salt that is formed.



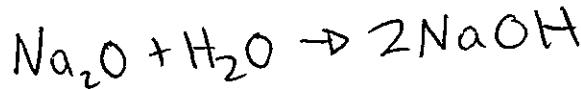
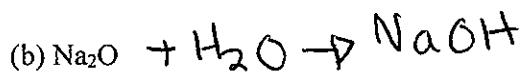
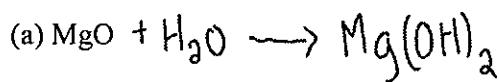
Metal oxides and water: when a metal oxide reacts with water to form a base

An oxide is a compound that contains oxygen and another element

Example of a metal oxide: potassium oxide (K_2O) or calcium oxide (CaO)



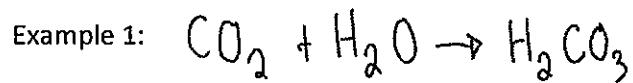
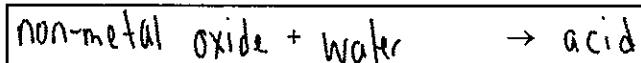
For each metal oxide, write the balanced equation when it is dissolved in water.



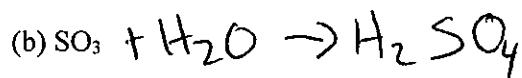
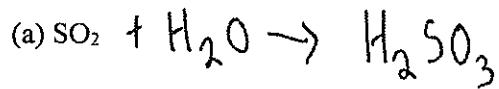
Non-metal Oxides and Water: when a non-metal oxide reacts with water to form an acid

A non-metal oxide is a compound that contains oxygen and another non-metal

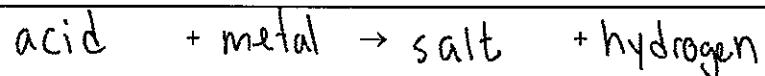
Example of a non-metal oxide: carbon dioxide (CO_2)



For each non-metal oxide, write the balanced equation when it is dissolved in water.



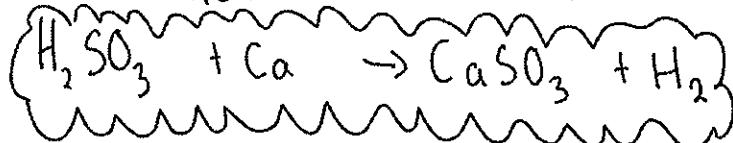
Acids and Metals: when an acid reacts with a metal to form salt and hydrogen



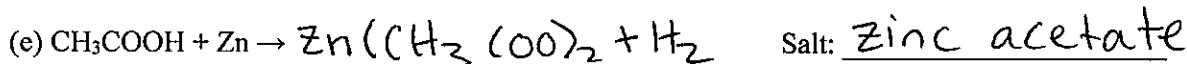
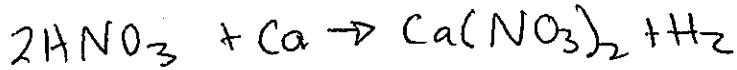
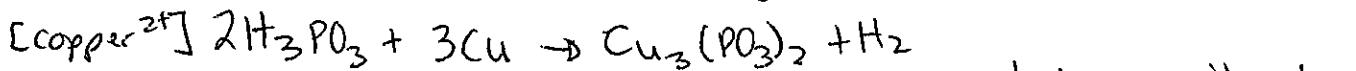
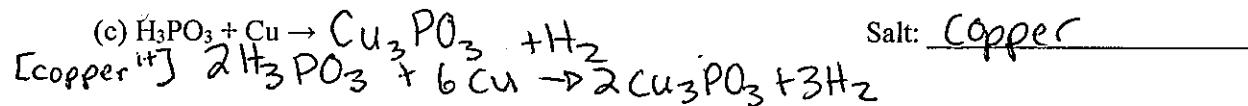
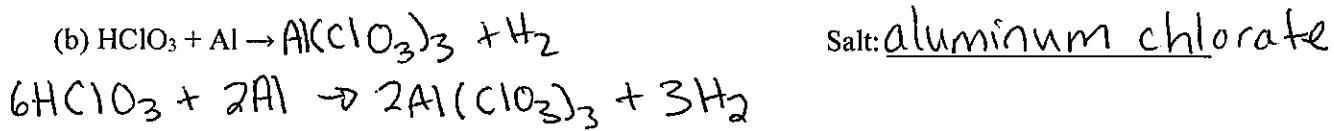
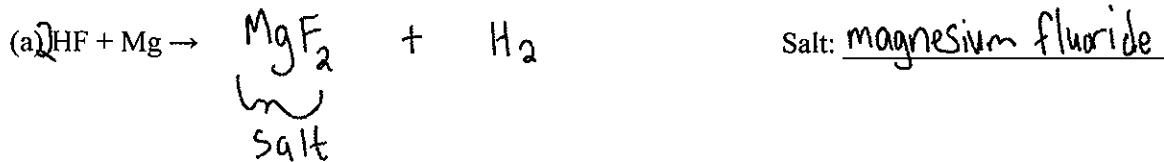
Example 1: hydrochloric acid + magnesium \rightarrow magnesium chloride + hydrogen



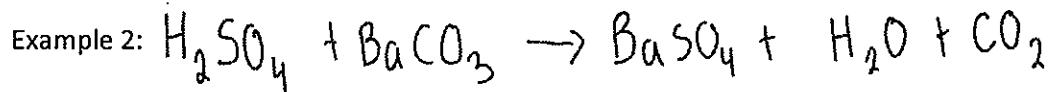
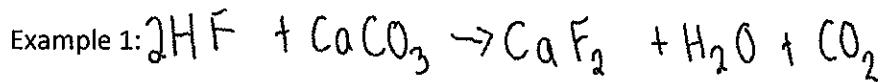
Example 2: sulfuric acid + calcium \rightarrow calcium sulfate + hydrogen



Complete and balance the following equations. Write the name of the salt.



Acids and Carbonates: when an acid reacts with a carbonate to form
a salt and water and carbon dioxide



Write the balanced equation for the following reactions:

