

Clé

Balancing Chemical Equations Worksheet

1.  $\underline{2} \text{H}_2 + \underline{1} \text{O}_2 \rightarrow \underline{2} \text{H}_2\text{O}$
2.  $\underline{1} \text{N}_2 + \underline{3} \text{H}_2 \rightarrow \underline{2} \text{NH}_3$
3.  $\underline{1} \text{S}_8 + \underline{12} \text{O}_2 \rightarrow \underline{8} \text{SO}_3$
4.  $\underline{2} \text{N}_2 + \underline{1} \text{O}_2 \rightarrow \underline{2} \text{N}_2\text{O}$
5.  $\underline{2} \text{HgO} \rightarrow \underline{2} \text{Hg} + \underline{1} \text{O}_2$
6.  $\underline{6} \text{CO}_2 + \underline{6} \text{H}_2\text{O} \rightarrow \underline{1} \text{C}_6\text{H}_{12}\text{O}_6 + \underline{6} \text{O}_2$
7.  $\underline{2} \text{Zn} + \underline{2} \text{HCl} \rightarrow \underline{1} \text{ZnCl}_2 + \underline{1} \text{H}_2$
8.  $\underline{1} \text{SiCl}_4 + \underline{4} \text{H}_2\text{O} \rightarrow \underline{1} \text{H}_4\text{SiO}_4 + \underline{4} \text{HCl}$
9.  $\underline{2} \text{Na} + \underline{2} \text{H}_2\text{O} \rightarrow \underline{2} \text{NaOH} + \underline{1} \text{H}_2$
10.  $\underline{2} \text{H}_3\text{PO}_4 \rightarrow \underline{1} \text{H}_4\text{P}_2\text{O}_7 + \underline{1} \text{H}_2\text{O}$
11.  $\underline{1} \text{C}_{10}\text{H}_{16} + \underline{8} \text{Cl}_2 \rightarrow \underline{10} \text{C} + \underline{16} \text{HCl}$
12.  $\underline{1} \text{CO}_2 + \underline{2} \text{NH}_3 \rightarrow \underline{1} \text{OC}(\text{NH}_2)_2 + \underline{1} \text{H}_2\text{O}$
13.  $\underline{4} \text{Si}_2\text{H}_3 + \underline{17} \text{O}_2 \rightarrow \underline{8} \text{SiO}_2 + \underline{6} \text{H}_2\text{O}_3$
14.  $\underline{2} \text{Al}(\text{OH})_3 + \underline{3} \text{H}_2\text{SO}_4 \rightarrow \underline{2} \text{Al}_2(\text{SO}_4)_3 + \underline{6} \text{H}_2\text{O}$
15.  $\underline{4} \text{Fe} + \underline{3} \text{O}_2 \rightarrow \underline{2} \text{Fe}_2\text{O}_3$
16.  $\underline{1} \text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
17.  $\underline{2} \text{C}_7\text{H}_6\text{O}_2 + \underline{15} \text{O}_2 \rightarrow \underline{14} \text{CO}_2 + \underline{6} \text{H}_2\text{O}$
18.  $\underline{1} \text{H}_2\text{SO}_4 + \underline{4} \text{HI} \rightarrow \underline{1} \text{H}_2\text{S} + \underline{2} \text{I}_2 + \underline{4} \text{H}_2\text{O}$
19.  $\underline{4} \text{FeS}_2 + \underline{11} \text{O}_2 \rightarrow \underline{2} \text{Fe}_2\text{O}_3 + \underline{8} \text{SO}_2$
20.  $\underline{2} \text{Al} + \underline{3} \text{FeO} \rightarrow \underline{1} \text{Al}_2\text{O}_3 + \underline{3} \text{Fe}$
21.  $\underline{1} \text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \underline{2} \text{Fe} + \underline{3} \text{H}_2\text{O}$
22.  $\underline{1} \text{Na}_2\text{CO}_3 + \underline{2} \text{HCl} \rightarrow \underline{2} \text{NaCl} + \underline{1} \text{H}_2\text{O} + \underline{1} \text{CO}_2$
23.  $\underline{2} \text{K} + \underline{1} \text{Br}_2 \rightarrow \underline{2} \text{KBr}$
24.  $\underline{1} \text{C}_7\text{H}_{16} + \underline{11} \text{O}_2 \rightarrow \underline{7} \text{CO}_2 + \underline{8} \text{H}_2\text{O}$
25.  $\underline{1} \text{P}_4 + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$